**实验11 金属与酸反应、金属与盐溶液反应——学习任务单**

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| 【实验目的】 探究金属的活动性  【实验过程】   |  |  |  | | --- | --- | --- | | 实验 | 实验现象 | 化学方程式 | | 镁与盐酸 | \_\_\_\_\_\_\_\_\_\_\_\_ | Mg + 2HCl= MgCl2 + H2↑ | | 镁与稀硫酸 | Mg + H2SO4 = MgSO4 + H2↑ | | 铝与盐酸 | \_\_\_\_\_\_\_\_\_\_\_\_ | 2Al + 6HCl= 2AlCl3 + 3H2↑ | | 铝与稀硫酸 | 2Al+ 3H2SO4 = Al2(SO4）3 + 3H2↑ | | 锌与盐酸 | \_\_\_\_\_\_\_\_\_\_\_\_ | Zn + 2HCl= ZnCl2 + H2↑ | | 锌与稀硫酸 | Zn + H2SO4 = Zn SO4 + H2↑ | | 铁与盐酸 | \_\_\_\_\_\_\_\_\_\_\_\_ | Fe + 2HCl= FeCl2 + H2↑ | | 铁与稀硫酸 | Fe + H2SO4 = FeSO4 + H2↑ | | 铜与盐酸 | \_\_\_\_\_\_\_\_\_\_\_\_ |  | | 铜与稀硫酸 |  | | 铝丝浸入硫酸铜溶液 | \_\_\_\_\_\_\_\_\_\_\_\_ | 2Al + 3CuSO4 = Al2(SO4)3 + 3Cu | | 铜丝浸入硝酸银溶液 | \_\_\_\_\_\_\_\_\_\_\_\_ | Cu + 2AgNO3 = Cu(NO3)2 + 2Ag | | 铜丝浸入硫酸铝溶液 | \_\_\_\_\_\_\_\_\_\_\_\_ |  | |
| 【实验结论】 金属活动性：\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  【反思拓展】（1）当其它条件均相同时，镁、铝、锌、铁与盐酸的反应速率由快到慢的顺序依次是\_\_\_\_\_\_\_\_。  （2）等质量的镁、铝、锌、铁与足量盐酸充分反应，产生氢气的质量由多到少依次是\_\_\_\_\_\_\_\_\_。 |

可以参考教材：九下 第八单元课题2 探究 P10